

Practice: Isotope Calculations #2

1. Complete the following table.

Isotopic Symbol (Nuclear form)	Isotope Symbol (Hyphen form)	Atomic #	Mass #	# of Protons	# of Neutrons	# of Electrons
			24	11		
			32	16		
					24	20
^{24}Mg						
			64		34	
	Lead-207					
		53	127			
^{238}U						
			129	52		
		9	19			

2. Explain the difference between mass number and average atomic mass.

3. Which tells you the most common version of an element?

4. Place a star next to each column that represents an isotope that is the most common for that element.